

## Gold(III)

Reaction	Baes and Mesmer, 1976
$\text{Au(OH)}_3 + 2 \text{H}^+ \rightleftharpoons \text{AuOH}^{2+} + 2 \text{H}_2\text{O}$	$1.51 \pm 0.2$
$\text{Au(OH)}_3 + \text{H}^+ \rightleftharpoons \text{Au(OH)}_2^+ + \text{H}_2\text{O}$	$< 1.0$
$\text{Au(OH)}_3 + \text{H}_2\text{O} \rightleftharpoons \text{Au(OH)}_4^- + \text{H}^+$	$-11.77 \pm 0.2$
$\text{Au(OH)}_3 + 2 \text{H}_2\text{O} \rightleftharpoons \text{Au(OH)}_5^{2-} + 2 \text{H}^+$	$-25.13 \pm 0.1$
$\text{Au(OH)}_5^{2-} + 3 \text{H}_2\text{O} \rightleftharpoons \text{Au(OH)}_6^{3-} + 3 \text{H}^+$	$< -41.1$
$\text{Au(OH)}_3(\text{c}) \rightleftharpoons \text{Au(OH)}_3$	$-5.51 \pm 0.07$