

Beryllium

Reaction	Baes and Mesmer, 1976
$\text{Be}^{2+} + \text{H}_2\text{O} \rightleftharpoons \text{BeOH}^+ + \text{H}^+$	-5.40
$\text{Be}^{2+} + 2 \text{H}_2\text{O} \rightleftharpoons \text{Be}(\text{OH})_2 + 2 \text{H}^+$	-23.65
$\text{Be}^{2+} + 3 \text{H}_2\text{O} \rightleftharpoons \text{Be}(\text{OH})_3^- + 3 \text{H}^+$	-23.25
$\text{Be}^{2+} + 4 \text{H}_2\text{O} \rightleftharpoons \text{Be}(\text{OH})_4^{2-} + 4 \text{H}^+$	-37.42
$2 \text{Be}^{2+} + \text{H}_2\text{O} \rightleftharpoons \text{Be}_2\text{OH}^{3+} + \text{H}^+$	-3.97
$3 \text{Be}^{2+} + 3 \text{H}_2\text{O} \rightleftharpoons \text{Be}_3(\text{OH})_3^{3+} + 3 \text{H}^+$	-8.92
$6 \text{Be}^{2+} + 8 \text{H}_2\text{O} \rightleftharpoons \text{Be}_6(\text{OH})_8^{4+} + 8 \text{H}^+$	-27.2
$\alpha\text{-Be}(\text{OH})_2(\text{cr}) + 2 \text{H}^+ \rightleftharpoons \text{Be}^{2+} + 2 \text{H}_2\text{O}$	6.69

C.F. Baes and R.E. Mesmer, The Hydrolysis of Cations. Wiley, New York, 1976.