

Tin(IV)

Reaction	Hummel et al., 2002	Gamsjäger et al, 2012	Brown and Ekberg, 2016
$\text{Sn}^{4+} + 4 \text{H}_2\text{O} = \text{Sn}(\text{OH})_4 + 4 \text{H}^+$			7.53 ± 0.12
$\text{Sn}^{4+} + 5 \text{H}_2\text{O} = \text{Sn}(\text{OH})_5^- + 5 \text{H}^+$			-1.07 ± 0.42
$\text{Sn}^{4+} + 6 \text{H}_2\text{O} = \text{Sn}(\text{OH})_6^{2-} + 6 \text{H}^+$			-11.14 ± 0.32
$\text{Sn}(\text{OH})_4 + \text{H}_2\text{O} = \text{Sn}(\text{OH})_5^- + \text{H}^+$	-8.0 ± 0.3	-8.60 ± 0.40	
$\text{Sn}(\text{OH})_4 + 2 \text{H}_2\text{O} = \text{Sn}(\text{OH})_6^{2-} + 2 \text{H}^+$	-18.4 ± 0.3	-18.67 ± 0.30	
$\text{SnO}_2(\text{cr}) + 2 \text{H}_2\text{O} = \text{Sn}(\text{OH})_4$	-8.0 ± 0.2	-8.06 ± 0.11	
$\text{SnO}_2(\text{am}) + 2 \text{H}_2\text{O} = \text{Sn}(\text{OH})_4$	-7.3 ± 0.3	-7.22 ± 0.08	

$\text{SnO}_2(\text{s}) + 4 \text{H}^+ = \text{Sn}^{4+} + 2 \text{H}_2\text{O}$			-15.59 ± 0.04
---	--	--	-------------------

P.L. Brown and C. Ekberg, Hydrolysis of Metal Ions. Wiley, 2016, pp. 836-842.

W. Hummel, U. Berner, E. Curti, F.J. Pearson and T. Thoenen. Nagra / PSI Chemical Thermodynamic Data Base 01/01, July 2002.

H. Gamsjäger, T. Gajda, J. Sangster, S. K. Saxena and W. Voigt. Chemical Thermodynamics of Tin. Chemical Thermodynamics Volume 12. OECD, Paris, 2012.